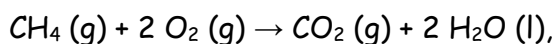
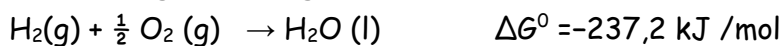


14.-Kalkulatu metanoaren errektuntza-erreakzioaren energia aske estandarra,



honako datu haueetatik abiatuta:



Emaitza: -818 kJ

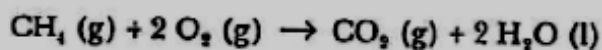
DATOAK FORMAZIO-ENERGIA ASKE MOLARRAK ETA ESTANDARRAK DIRENEZ, HESSEN LEGEAREN ONDORIOZ SORTZEN DEN FORMULA, ERREAKZIOAN, APLIKATUKO DUGU.

Datuak: a) $\Delta G^\circ_f [\text{CH}_4(\text{g})] = -50,8 \text{ kJ/mol}$

b) $\Delta G^\circ_f [\text{CO}_2(\text{g})] = -394,4 \text{ kJ/mol}$

c) $\Delta G^\circ_f [\text{H}_2\text{O}(\text{l})] = -237,2 \text{ kJ/mol}$

Erreakzioa:



$$\Delta G^\circ_r = \sum n \Delta G^\circ_{f \text{ produktuak}} - \sum m \Delta G^\circ_{f \text{ erreaktibok}}$$

$$\Delta G^\circ_r = 2 \Delta G^\circ_f [\text{H}_2\text{O}(\text{l})] - 1 \Delta G^\circ_f [\text{CO}_2(\text{g})] - 1 \Delta G^\circ_f [\text{CH}_4(\text{g})]$$

UNITATEAK: mol. KJ/mol = KJ

$$\Delta G^\circ_r = 2 \cdot (-237,2 \text{ kJ}) + 1 \cdot (-394,4 \text{ kJ}) - 1 \cdot (-50,8 \text{ kJ})$$

$$\Delta G^\circ_r = -818,0 \text{ kJ}$$

Erreakzioaren energia aske estandarra - 818 kJ-koa da.